

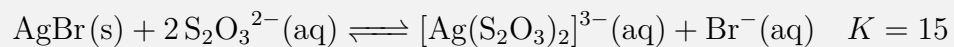
- (b) All the radon gas is removed from the container and replaced by air. The system is re-established at equilibrium at 298 K. What is the partial pressure of radon in the vessel at this new equilibrium? A picture can help understand what physical process is occurring.

3. A solid has a density of 5.3 g/cm^3 while its liquid phase has a density of 5.6 g/cm^3 . Circle the statements that describe the change(s) that the substance would undergo in response to a decrease in external pressure.

- The freezing point of the liquid will decrease
- The melting point of the solid will increase
- The boiling point of the liquid will increase
- The condensation point of the gas will decrease
- The intermolecular forces of the liquid phase will increase

Homework Problem 27

1.



You want to dissolve 15 g of AgBr by adding sodium thiosulfate pentahydrate ($\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}$) to water. What is the minimum mass of $\text{Na}_2\text{S}_2\text{O}_3 \cdot 5\text{H}_2\text{O}(s)$ that must be added to 500 mL of water in order to ensure that all 15 g of AgBr will dissolve at equilibrium?